The Crystal Structure of $[Pt(CH_3)_3(rac-diars)]I;$ (diars= $C_6H_5CH_3AsCH_2CH_3AsCH_3C_6H_5$)

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Crystals of $Pt(CH_3)_3(C_6H_5CH_3AsCH_2CH_2AsCH_3C_6H_5)I$ are monoclinic with a = 17.008, b = 15.736, c = 8.393 Å, β = 101.37°, space group $P2_1/c$, Z=4. The definition of atomic parameters was by Fourier and least-squares analysis of 1976 independent reflections (MoKa, four circle diffractometry); convergence was to $R_1 = 0.06$ for the absorption corrected data.

The mean Pt-C(methyl) bond length is 2.11 Å, the Pt-As and Pt-I bond lengths being 2.455 and 2.771 Å The trans-influence of the methyl group on the Pt-As and Pt-I bonds is shown to be ca. 0.1 Å; the conformation of the chelate ring is a distorted envelope.

Introduction

During the course of their systematic studies of transitions metal-carbon bonds, Shaw and co-workers¹

have described oxidative addition reactions of $[PtMe_2(rac-diars)]$ or $[PtMe_2(meso-diars)]$, $(diars = C_6H_5CH_3-AsCH_2CH_2AsCH_3C_6H_5)$; the complexes add methyl iodide stereospecifically but it was not initially clear what the conformation of the coordinated arsine was in the two adducts. An X-ray analysis provided clear information on the point but also allows some interesting conclusions on *cis*- and *trans*-influence in six coordinate complexes and on general conformational aspects of unsymmetrical coordinated diarsines.

Experimental Section

Pale yellow prisms from benzene-light petroleum were shown to be monoclinic by Weissenberg and precession photographs (MoK α). A selected crystal of dimension $0.1 \times 0.4 \times 0.3$ mm was used for intensity

Table I. $Pt(CH_3)_3(C_6H_5CH_3A_5CH_2CH_2A_5CH_3C_6H_3)I.$

(i)]	Heavy-atom parameters	and their e.s.c	's (multiplied	by 10 ⁴). T	he temperatu	re factor is	in the form	1	
		exp	$-(\beta_{11}h^2 + \beta_{22}k^2 +$	$-\beta_{33}l^2+2\beta_{12}h$	$k+2\beta_{13}hl+2\beta$	23kl)]			
	х	У	z	β.	β22	β.,	β12	βι	β ₂₃
Pt	2873(1)	237(1)	3989(1)	37(0)	41(0)	113(1)	2(0)	11(1)	-2(1)
I	1638(1)	-336(1)	5428(2)	64(1)	67(1)	193(3)	$\hat{0(1)}$	59(1)	18(2)
As(1)	2800(2)	-1087(1)	2407(3)	41(1)	31(1)	127(4)	1(1)	19(2)	7(2)
As(2)	1887(1)	729(1)	1637(3)	40(1)	34(1)	117(4)	0(1)	15(2)	8(2)
(ii)	Carbon-atom parameter	rs and their e.	s.d's.						
		x		у			z		В
C(1)	0	.2904(19)		0.1406(2	Ð	0.5	255(38)		5.1(7)
C(2)	0	3749(19)		0.0270(2	ĺ)	0.5	919(38)		6.8(7)
C(3)	0	.3903(17)		0.0703(1	8)	0.3	276(34)		5.5(6)
C(4)	0	.2535(18)		0.2151(1	8)	0.3	425(35)		6,7(6)
C(5)	0	.2327(16)		0.1166(1	7)	0.02	282(31)		5.6(5)
C(6)	0.	.1948(16)		0.0998(1	6)	0.04	482(31)		5.0(5)
C(7)	0	.1350(14)		0.0301(1	5)	0.0	652(27)		4.6(4)
C(8)	0	.3758(14)		-0.1360(1	5)	0.1	577(28)		4.1(5)
C(9)	0.	.3945(15)		0.0973(1	5)	0.0	267(29)		3.8(5)
C(10)	0.	.4682(17)		0.1096(1	9)	0.0	199(34)		5.7(6)
C(11)	0.	.5214(19)		-0.1617(1	9)	0.0	574(37)		6.0(6)
C(12)	0	.5027(21)		0.2039(2	1)	0.1	914(40)		7.2(8)
C(13)	0.	4329(17)			7)	0.24	481(33)		5.6(6)
C(14)	0.	.1038(14)		0.1518(1	4)	0.1	798(27)		4.0(4)
C(15)	0.	.0357(16)		0.1246(1	6)	0.2	356(31)		5.6(5)
C(16)	0.	.0232(17)		0.1849(1	8)	0.2	524(34)		6.2(6)
C(17)	0.	0152(17)		0.2708(1)	3)	0.2	77(34)		5.8(6)
C(18)	0.	0544(16)		0.2962(1)	7)	0.15	570(32)		5.3(5)
C(19)	0	.1154(17)		0.2365(1)	7)	0.14	487(34)		5.9(6)

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K L FO FC	K L FO FC	K L F3 F5	CL FO FC	K L FO FC	K L FO	FG K L FG FC	< L F0 FC < L F0 FC

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Table II. (Continued)

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K L FU FC	к. с. на 1-7,443	FC 7342	K L FU	-294	KLF	U FC	K L	F0	FC -315	K L	FU 663	24 949-	K L	FU	FC -783	K L ął	FD	FC	к L 	FO	FC	
U -7 514 -535 9 -6 584 311 9 -8 316 164 9 -2 1207 -1379 9 -2 1977	1 5 731 2 6 696 2 0 909 2 3 576	-587 -688 704 1 -869 1 -869 1	9 -2 431 9 -2 715 9 -2 75 9 -	-713 -713 -858 -859 -870	-3195	2 -1483 918 6 438 2 473 473 1 409	н и - 8	14 630 499	641	0.0777	405 405 405 514 514	587 -539 -420 -420 -404 541		470 471 632 810 400	- 446 661 - 016 853		854 655 574 57 574 57			4123 710 3810 3810	-71909	
10 -1 793 801 10 -1 793 800 800 800 800 800 800 800 800 800 80	3 -5 980 3 -1 450 3 -1 473 3 -1 473 3 -1 473 3 -5 480 3 -5 -	-1015 L +451 L +611 L -646 -846 -480		-315		77 25 -1042 -1042 -591 -591 -591 -591	1	485 556 395 395 	-575	8 -7 U -5	5424U73	-541	nn 0 0 0 7 7	001633 0733045	-373 -312 -706 -813 -440 -440 -440	H	407	-239 -639 -382		4043	-386 -549 -509 484 128 -371	
LI -2 1148 -1180 LI 0 473 -429 L2 0 473 417 12 -1 954 533 L2 1 646 675 12 3 620 -608	1 105 2 779 - 3 906 - 833 - 835 - 8 537 - 8 535	1035 -440 654 776 -360 -544	J -4 1082 J -2 1955 J 0 440 1 -7 481		7 -6 87 7 -4 49 7 -3 74 7 110 7 156 8 -7 48	5 -867 -439 1 -816 3 1130 8 576 599		620 504 748 552 599	552 -455 -655 -655 -555 -555		406 618 787 417	-480 489 008 -269		5475 711 541	442 418 -710 -668 -505	1000	795 554 533 582 407	-502 -559 -597 -597	5 -1 7 -1	\$73 \$33 \$33	-594 492 -290 624	
13		-64774	1 -1 1302 1 -1 1302 1 -2 1302 -3 657 -3 657 -7 1569	-L365 951 13779	8 -5 61 -5 88 -5 729 -6 88 -7 20 -6 88 -7 20 -7 20	8 -895 -895 -7125 -715 -715 -715 -715 -715 -715 -715 -71		>28 575 56 55 66 57 56 56 56 56 56 56 56 56 56 56 56 56 56	2003D4	H	15 786 433 567	-1497 -391 -20 -1011		458 83951 37013 8498 84913 8498	5251 1051 10593 10593 10593		07504L87	-708	101-100	10		
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U -0 471 504 U -2 478 504 U -2 1109 -1292 U 4 1379 -1896	7 -8 670	- 426 039 - 341	4 -3 1508 3 -1 703 3 -2 862 8 -8 793 4 -8 433	-1521 -755 148 -107	-2 61 -2 61 -2 62 -2 62 -5 65	7 -93U 5 626 7 610 7 506 6 -627	****	587 407 476	-547 589 403 -536 377		432	1220 -554 -541 -541	0 1 - 1 - 1	7537 5521 5511	- 475		29562	-201 340 569 284 -362	î -î	514		
	7 5 462 8 -7 904 8 -3 436	- 4 38 920 - 448		-132		-523	5 -7	51 F	-544	3 3	613 755	->71	2 2 2	292	-639 -537 526	13	323	-428				

the Å

Table III. $Pt(CH_3)_3(C_6H_3CH_3A_5CH_2CH_2A_5CH_3C_6H_5)I$.

(i)	Bond lengths (Å) and	their e.s.d's.	
	Platinum-ligand bonds:		
	Pt-I	2.771(2)	
	Pt-As(1)	2.461(3)	
	Pt-As(2)	2.450(3)	
	Pt-C(1)	2.12(3)	
	Pt-C(2)	2.13(3)	
	Pt-C(3)	2,09(3)	
	Arsenic-carbon bonds:		
	As(1)-C(4)	1.97(3)	
	As(1)-C(6)	1.95(3)	
	As(1)-C(8)	1.94(2)	
	As(2)-C(5)	2.02(3)	
	As(2)-C(7)	1.96(2)	
	As(2)-C(14)	1.93(2)	
	Carbon-carbon bonds:		
	C(6)-C(7)	1.52(4)	
	C(8)-C(9)	1.35(3)	
	C(9)-C(10)	1.40(3)	
	C(10)-C11)	1.33(4)	
	C(11)-C(12)	1.32(5)	
	C(12)-C(13)	1.38(5)	Mean value of
	C(13)-C(8)	1.42(4)	(C-C) _{penyl}
	C(14)-C(15)	1.40(4)	distance $= 1.39$
	C(15)-C(16)	1.41(4)	
	C(16)-C(17)	1.40(4)	
	C(17)-C(18)	1.43(4)	
	C(18)-C(19)	1.41(4)	
	C(19)-C(14)	1.38(4)	
(ii)	Bond angles (degrees)	and their e.	s.d's.
	I-Pt-As(1)	89.9(1)	
	I-Pt-As(2)	89.9(1)	
	I-Pt-C(1)	90.6(9)	
	I-Pt-C(2)	91.4(9)	
	I-Pt-C(3)	171.0(8)	
	As(1)-Pt- $As(2)$	83.1(1)	
	As(1)-Pt-C(1)	177.3(8)	
	As(1)-Pt-C(2)	92.7(9)	
	As(1)-Pt-C(3)	95.6(8)	
	As(2)-Pt-C(1)	94.3(9)	
	As(2)-Pt-C(2)	175.6(9)	
	As(2)-Pt-C(3)	97.9(8)	
	C(1)-Pt-C(2)	90.0(12)	
	C(1)-Pt-C(3)	84.3(12)	
	C(2)-Pt-C(3)	81.3(12)	
	Pt-As(1)-C(4)	118.5(9)	
	Pt-As(1)-C(6)	109.6(8)	
	Pt-As(1)-C(8)	115.1(7)	

Pt-As(2)-C(5)	116.5(8)
Pt-As(2)-C(7)	105.4(7)
Pt-As(2)-C(14)	123.3(7)
C(4)-As(1)-C(6)	103.0(12)
C(4)-As(1)-C(8)	104.3(11)

Table III. (continued)

C(6)-As(1)-C(8)

C(5)-As(2)-C(7)

C(5)-As(2)-C(14)

C(7)-As(2)-C(14)

As(1)-C(6)-C(7)

As(2)-C(7)-C(6)	111.8(17)		
(C-C-C) _{peny1}	120.0(30)	mean value	
measurements, the un	it cell parame	ters being d	letermi-
ned from four circle	setting of 12	reflections	(MoKa
$\lambda = 0.7107$ Å). The u	init cell is a	= 17.008(6), $b =$
15.736(5), c = 8.39	$3(3)$ Å, $\beta =$	= 101.37(3) ⁶	;* For
$PtIAs_2C_{19}H_{29}: M =$	729.3, U =	2202 Å ³ ,	$D_m =$
2.183 g. cm ⁻³ (flotatio	on), $Z = 4, D$	c = 2.190 g	. cm ⁻³ ,
F(000) = 1360. T	he absorptior	coefficient	μ(Mo-
$K\alpha$) = 111.7 cm ⁻¹ .	The system	atic absence	es {h0l}
for $1 = 2n+1$ and	{0k0} for k	= 2n+1 e	stablish
the space group $P2_1/$	c.		

104.9(11)

99.2(10)

103.7(10)

105.6(10)

112.1(17)

Intensity measurements were based on an ω -2 θ scan routine, background measurements at each extremity of a scan being made with stationary crystal and counter. Each background count was over 25 or 50 seconds, each peak count being accumulated over 50 steps of 0.01° with 1 (or 2) second counts respectively at each step.

Of the 3876 independent reflections measured (4° \leq $2\theta \leq 50^{\circ}$), 1998 were judged to be significant (I \geq 3.0. $\sigma(I)$). These were corrected for geometrical effects and, using the method of Wells, for absorption. Stability of the crystal and of the instrument was monitored by following the intensities of 3 reference reflections every 100 reflections.

The platinum, iodine and arsenic atoms were obtained, after some difficulties, from a Patterson synthesis after which the structure determination was based on a straightforward application of the difference Fourier method. Initial refinement of atomic para-

(*) Here as elsewhere, the figures in parentheses refer to the e.s.d's of the observation referred to the last significant figure.

meters was via block-diagonal least-squares methods but refinement of atomic positions, anisotropic temperature factors for the platinum, iodine and arsenic atoms and isotropic temperature factors for the remaining non-hydrogen atoms was completed by three full-matrix least-squares cycles. 22 reflections, apparently suffering from extinction effects, were excluded from these terminal cycles and convergence was to $R_1 = 0.06$ and $R_2 = 0.08$. The weighting scheme was w = $(55.03-0.61 \cdot |F_{obs}| + 0.005 \cdot |F_{obs}|^2)^{-1}$ and was obtained from a least-squares fit of $(|F_{obs}| - |F_{cal}|)^2$ as a function of | Fobs |. In the structure factor calculations, the atomic scattering factors of Cromer and Waber² were used for platinum, iodine, arsenic and carbon atoms.

Atomic parameters and their e.s.d's are listed in Table I, observed and calculated structure factors being collected in Table II. Atom numbering and some important intramolecular bond lengths and bond angles are shown in Figure 1, a more complete listing of molecular parameters forming Table III.



Figure 1. $Pt(CH_3)_3(C_6H_5CH_3A_5CH_2CH_2A_5CH_3C_6H_5)I$. Numbering of atoms and some bond lengths and bond angles. The molecule is viewed along the c axis.

Discussion

Two general features of the molecular geometry deserve emphasis.

(i) The metal-ligand bond lengths. There seem to be remarkably few (if any) really accurate measurements of the platinum(IV)-iodine bond length. In diiodo-di-(o-phenylenebis (dimethylarsine)) platinum(II)³ the platinum-arsenic and platinum-iodine bond lengths are 2.38Å and 3.50 Å respectively; the latter value is obviously not comparable with the present

result, the bond being essentially "ionic" in the platinum(II) complex. The covalent radii of arsenic, platinum(II) and iodine atoms are ca. 1.20, 1.30 and 1.33 Å respectively; the platinum-carbon bond lengths are thus close to the sum of covalent radii (2.07 Å) and the mean value of 2.11 Å is not significantly different from those found in acetylacetonyl-2,2'-bipyridyltrimethylplatinum(IV)⁴ (2.05 Å), in salicylaldehydatotrimethylplatinum(IV)⁵ (2.03 Å), or indeed, the platinum(II)-carbon bond length of 2.11 Å in the bisacetylacetonatochloroplatinate(II) anion.6 That the platinum-arsenic bond length is also close to the sum of covalent radii is fortuitous; it is a general observation that the metal-arsenic and metal-phosphorus bond lengths are normally 0.10-0.15 Å shorter than the sum of covalent radii in the absence of strong trans-influencing ligands. Whatever the reason for the "normal" foreshortening of metal-arsenic bonds, it is offset in the present structure by the lengthening of the bonds trans to methyl groups. Elsewhere we have estimated the trans-influence of a σ -bonded carbon atom in planar and octahedral complexes as approximately 0.1 Å. This is essentially the difference between the sum of the platinum and iodine covalent radii (2.63) Å) and the observed value.

(ii) The chelate ring geometry and the conformation of substituents. The chelate ring has a distorted envelope conformation. The mean plane containing Pt, As(1), As(2), C(1) and C(2) atoms has direction cosines 0.8376, 0.3252, -0.5955 referred to the crystallographic axes a, b, and c; the r.m.s. deviation of these atoms from the plane is 0.01 Å. The C(6) atom is displaced -0.21 Å from the plane, C(7) being -0.79 Å (both are bent towards the iodine atom). The two methyl carbon atoms C(4) and C(5) are displaced -1.43 Å and 1.82 Å respectively, the phenyl carbon atoms C(8) and C(14) having equivalent displacements of 1.64 Å and -0.88 Å so that the dl form of the complex is established as (1), the meso form as (11):



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